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of Mount St. Mary ' s College in Los Angeles, California. They give an overview of each free-response question and of how students performed on the question, including typical student errors.

AP Chemistry 2009 Q&A - College Board

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$$\text{CH}_4(\text{g}) + 2 \text{Cl}_2(\text{g}) \rightarrow \text{CH}_2\text{Cl}_2(\text{g}) + 2 \text{HCl}(\text{g})$$

3. Methane gas reacts with chlorine gas to form dichloromethane and hydrogen chloride, as represented by the equation above. (a) A 25.0 g sample of methane gas is placed in a reaction vessel containing 2.58 mol of $\text{Cl}_2(\text{g})$.

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If the temperature is increased for
an endothermic

reaction($\Delta H_{298}^{\circ} = +131 \text{ kJ}$
 mol^{-1}), then by Le Chatelier ' s
principle the reaction will shift
toward products, thereby
absorbing energy. With greater
concentrations of products at
equilibrium, the value of K_p will
increase.

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Answer Questions 1, 2, and 3.

Answer Questions 1, 2, and 3. The Section II score weighting for each question is 20 percent.

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and G), writing a K_p expression
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Question 2009 #5

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