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$h = h_{vap} = \exp \left(\frac{14.790}{T} - \frac{5301.5}{T^2} \right)$ and by the ideal gas law $h = \frac{RT}{P} \ln \left(\frac{P}{P^*} \right) + h_{vap}$ Solutions to Chemical and Engineering Thermodynamics, 3e $N = \frac{P}{RT} \times 18 \text{ mol} \times 8.314 \times 10^{-5} \text{ (bar} \cdot \text{m}^3 \text{ mol}^{-1} \text{K}^{-1})$ $T P = \frac{V}{V_1} \times 10^{-3} - (10 - x) 106$ volume taken up by liquid Equating P and P vap we have $x = 18 \times 8.314 \times 10^{-5} T F 5432.8 \times I$.

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